

## More with Scientific Notation – NOTES

Name: \_\_\_\_\_

Date: \_\_\_\_\_

### Multiplying

- To multiply with numbers in scientific notation:

1. \_\_\_\_\_

2. \_\_\_\_\_

- Examples:

1.  $(3.5 \times 10^{12}) * (6.7 \times 10^3)$

2.  $(5.2 \times 10^9) * (3.81 \times 10^2)$

3.  $(4.05 \times 10^{-3}) * (5.202 \times 10^{-5})$

4.  $(6.3 \times 10^{-10}) * (7.0 \times 10^4)$

5. There are  $2.7 \times 10^8$  hemoglobin molecules in a single human red blood cell. There are about  $5.0 \times 10^6$  human red blood cells in one small drop of blood. How many hemoglobin molecules are in one small drop of blood?

## Dividing

- To divide with numbers in scientific notation:

1. \_\_\_\_\_

2. \_\_\_\_\_

- Examples:

1.  $(3.5 \times 10^{12}) \div (6.7 \times 10^3)$

2.  $(5.2 \times 10^9) \div (3.81 \times 10^2)$

3.  $(4.05 \times 10^{-3}) \div (5.202 \times 10^{-5})$

4.  $(6.3 \times 10^{-10}) \div (7.0 \times 10^4)$

5. Suppose you have a gold nugget that has a mass of 2.41 grams. There are  $7.4 \times 10^{21}$  atoms of gold in the nugget. What is the mass of one atom of gold?